vent,[14] such as pyridine, there is a strong increase in magnetic susceptibility with increasing pressure (Table I). This effect is readily explained in terms of the ionization equilibrium

$$Cu_2A_4 + \text{solvent} \overset{P}{\rightleftharpoons} 2Cu^{2+}(\text{solvent}) + 4A^-(\text{solvent}) \quad (1)$$

where A represents an alkanoate. In pure water copper acetate appears to be almost completely ionized. In agreement with other results,[1,21,22] increasing pressure shifts the equilibrium in favor of the more solvated species. Equilibrium 1 parallels the pressure-dependent equilibria between $CuCl_4^{2-}$ or $CoCl_4^{2-}$ and the solvated Cu^{2+} and Cl^- or Co^{2+} and Cl^- species.[21]

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